

Unit 7 organic chemistry			
1	Hydrocarbon	molecules made up of hydrogen and carbon only	
2	Crude oil	a mixture of different length hydrocarbons formed over millions of years from dead organisms	
3	Alkane	saturated hydrocarbon containing hydrogen and carbon only	
4	Saturated	no double bond	
5	Alkane formula	C_nH_{2n+2}	
6	First 4 alkanes	Alkane name	Number of carbons
		Methane	1
		Ethane	2
		Propane	3
	Butane	4	
7	Fractional distillation	a method of separating crude oil	
8	Fractional distillation process	evaporate then condense at different temperatures as each fraction has a different boiling points	
9	Boiling point	temperature a substance changes from a liquid to a gas	
10	Viscosity	how difficult it is for a substance to flow - the higher the viscosity, the 'thicker' it is	
11	Flammability	how easy it is to ignite and burn	
12	Properties of small chain hydrocarbons	low boiling point highly flammable low viscosity	
13	Properties of large chain hydrocarbons	high boiling point difficult to ignite high viscosity	
14	Cracking	breaking down large hydrocarbons to make more useful, shorter chains of an alkene and alkane	
15	Testing for alkenes	bromine water will change from orange to colourless when added to an alkene	

Unit 8 chemical analysis		
16	Pure substance	single element or compound, not mixed with any other substance
17	Pure substance properties	melt and boil at specific temperatures
18	Formulation	mixture that has been designed as a useful product
Required practical - Chromatography		
19	Chromatography	used to separate soluble substances
20	Stationary phase	chromatography paper
21	Mobile phase	the solvent
22	Rf value	the ratio of the distance moved by the solvent
23	Rf value calculation	distance moved by substance/ distance moved by solvent
24	Chromatography method	a. draw a pencil line on bottom of paper b. add spots of each sample to the line c. put in a container with solvent in the bottom d. allow the solvent to move through the paper e. measure the distance travelled by the spots and solvent
Gas tests		
25	Oxygen	relights a glowing splint
26	Hydrogen	burning splint will burn rapidly with a squeaky pop sound
27	Carbon dioxide	turns lime water milky
28	Chlorine	damp litmus paper will be bleached and turn white